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SUBMIT

Balancing Chemical Equations

Balance the equations below:

- 1) $1 \text{ N}_2 + 3 \text{ H}_2 \rightarrow 2 \text{ NH}_3$
- 2) $2 \text{ KClO}_3 \rightarrow 2 \text{ KCl} + 3 \text{ O}_2$
- 3) $2 \text{ NaCl} + 1 \text{ F}_2 \rightarrow 2 \text{ NaF} + 1 \text{ Cl}_2$
- 4) $2 \text{ H}_2 + 1 \text{ O}_2 \rightarrow 2 \text{ H}_2\text{O}$
- 5) $1 \text{ Pb(OH)}_2 + 2 \text{ HCl} \rightarrow 2 \text{ H}_2\text{O} + 1 \text{ PbCl}_2$
- 6) $2 \text{ AlBr}_3 + 3 \text{ K}_2\text{SO}_4 \rightarrow 6 \text{ KBr} + 1 \text{ Al}_2(\text{SO}_4)_3$
- 7) $1 \text{ CH}_4 + 2 \text{ O}_2 \rightarrow 1 \text{ CO}_2 + 2 \text{ H}_2\text{O}$
- 8) $1 \text{ C}_2\text{H}_6 + 5 \text{ O}_2 \rightarrow 3 \text{ CO}_2 + 4 \text{ H}_2\text{O}$
- 9) $2 \text{ C}_6\text{H}_{14} + 25 \text{ O}_2 \rightarrow 16 \text{ CO}_2 + 18 \text{ H}_2\text{O}$
- 10) $1 \text{ FeCl}_3 + 3 \text{ NaOH} \rightarrow 1 \text{ Fe(OH)}_3 + 3 \text{ NaCl}$
- 11) $4 \text{ P} + 5 \text{ O}_2 \rightarrow 2 \text{ P}_2\text{O}_5$
- 12) $2 \text{ Na} + 2 \text{ H}_2\text{O} \rightarrow 2 \text{ NaOH} + 1 \text{ H}_2$
- 13) $2 \text{ Ag}_2\text{O} \rightarrow 4 \text{ Ag} + 1 \text{ O}_2$
- 14) $1 \text{ S}_8 + 12 \text{ O}_2 \rightarrow 8 \text{ SO}_2$
- 15) $6 \text{ CO}_2 + 6 \text{ H}_2\text{O} \rightarrow 1 \text{ C}_6\text{H}_{12}\text{O}_6 + 6 \text{ O}_2$
- 16) $2 \text{ K} + 1 \text{ MgBr}_2 \rightarrow 2 \text{ KBr} + 1 \text{ Mg}$
- 17) $2 \text{ HCl} + 1 \text{ CaCO}_3 \rightarrow 1 \text{ CaCl}_2 + 1 \text{ H}_2\text{O} + 1 \text{ CO}_2$
- 18) $1 \text{ HNO}_3 + 1 \text{ NaHCO}_3 \rightarrow 1 \text{ NaNO}_3 + 1 \text{ H}_2\text{O} + 1 \text{ CO}_2$
- 19) $2 \text{ H}_2\text{O} + 1 \text{ O}_2 \rightarrow 2 \text{ H}_2\text{O}_2$
- 20) $2 \text{ NaBr} + 1 \text{ CaF}_2 \rightarrow 2 \text{ NaF} + 1 \text{ CaBr}_2$
- 21) $1 \text{ H}_2\text{SO}_4 + 2 \text{ NaNO}_2 \rightarrow 2 \text{ HNO}_2 + 1 \text{ Na}_2\text{SO}_4$

Handwritten notes:
 H₂O, Br₂
 ERROR
 H₂O, Br₂

l.) What volume of a 0.500 M MnCl₂ solution is needed to prepare 200.0 ml of a 0.150 M MnCl₂ solution?

$$(0.500 \text{ M})(V_1) = (0.150 \text{ M})(200.0 \text{ mL})$$

$$V_1 = 60.0 \text{ mL}$$

j.) 250.0 ml of a 0.0100 M solution of the salt mercury(II) nitrate sits out in an open beaker. After several days the volume of solution is reduced to 208.0 ml. What is the molarity of this new solution after evaporation. If this solution is rediluted to 300.0 ml, what is the new molarity after dilution? (2 answers)

$$(250.0 \text{ mL})(0.0100 \text{ M}) = M_2(208.0 \text{ mL})$$

$$M_2 = 0.0120 \text{ M}$$

$$(0.0100 \text{ M})(250.0 \text{ mL}) = M_2(300.0 \text{ mL})$$

$$M_2 = 0.00833 \text{ M}$$

k.) What mass of vitamin C (formula mass 176 g/mol) is needed to prepare a 450.0 ml of a 2.09 M vitamin C solution?

$$2.09 \text{ M} \cdot 0.4500 \text{ L} = 0.9405 \text{ mol} \cdot \frac{176 \text{ g}}{1 \text{ mol}} = 166 \text{ g}$$

l.) 6.00 ml of a 0.558 M ferrous gluconate solution is diluted to a new volume of 50.0 ml. What is the new molarity of the solution?

$$(0.558 \text{ M})(6.00 \text{ mL}) = (M_2)(50.0 \text{ mL})$$

$$M_2 = 0.0670 \text{ M}$$

m.) How many grams of potassium sulfate are needed to prepare 200.0 ml of a 0.125 M solution?

$$0.125 \text{ M} \cdot 0.2000 \text{ L} = 0.0250 \text{ mol} \cdot \frac{174 \text{ g}}{1 \text{ mol}} = 4.35 \text{ g K}_2\text{SO}_4$$

n.) Calculate the molar concentration of all ions in a solution that contains 82.31g of copper(II) nitrate in 912 ml of solution.

$$\text{Cu(NO}_3)_2 \rightarrow \text{Cu}^{2+} + 2\text{NO}_3^-$$

$$0.481 \text{ M} \rightarrow [0.481] \quad [0.963]$$

$$\frac{82.31 \text{ g}}{0.912 \text{ L}} \cdot \frac{1 \text{ mol}}{187.5 \text{ g}} = 0.481 \text{ M}$$

o.) A student places 2.00g of salt containing chromium in some water. Later it is tested for chromium ions, however there are no chromium ions in solution. What can be concluded about the chromium salt?

THE CR SALT IS INSOLUBLE SINCE IT DIDN'T DISSOCIATE INTO IONS.

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