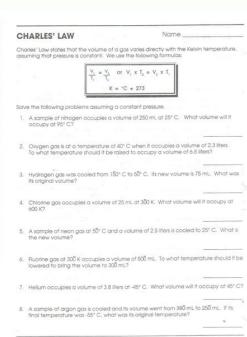




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Word equations worksheet chemistry answer key answer keys printable



Balancing Chemical Equations

Balance the equations below:

- 1) 1 N₂ + 3 H₂ → 2 NH₃

2) 2 KClO₃ → 2 KCl + 3 O₂

3) 2 NaCl + 1 F₂ → 2 NaF + 1 Cl₂

4) 2 H₂ + 1 O₂ → 2 H₂O

5) 1 Pb(OH)₂ + 2 HCl → 1 H₂O + 1 PbCl₂

6) 2 AlBr₃ + 3 K₂SO₄ → 6 KBr + 1 Al₂(SO₄)₃

7) 1 CH₄ + 2 O₂ → 1 CO₂ + 2 H₂O

8) 1 C₃H₈ + 5 O₂ → 3 CO₂ + 4 H₂O

9) 2 C₆H₁₆ + 12 O₂ → 16 CO₂ + 8 H₂O

10) 1 FeCl₃ + 3 NaOH → 1 Fe(OH)₃ + 3 NaCl

11) 4 P + 5 O₂ → 2 P₂O₅

12) 2 Na + 2 H₂O → 2 NaOH + 1 H₂

13) 2 Ag₂O → 4 Ag + 1 O₂

14) 1 S₈ + 12 O₂ → 8 SO₃

15) 6 CO₂ + 6 H₂O → 1 C₆H₁₂O₆ + 6 O₂

16) 2 K + 1 MgBr₂ → 2 KBr + 1 Mg

17) 2 HCl + 1 CaCO₃ → 1 CaCl₂ + 1 H₂O + 1 CO₂

18) 1 HNO₃ + 1 NaHCO₃ → 1 NaNO₃ + 1 H₂O + 1 CO₂

19) 2 H₂O + 1 O₂ → 2 H₂O₂

20) 2 NaBr + 1 CaF₂ → 2 NaF + 1 CaBr₂

21) 1 H₂SO₄ + 2 NaNO₃ → 2 HNO₃ + 1 Na₂SO₄

For example both *soft* and *new* phenomena are

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- i.) What volume of a 0.500 M MnCl_2 solution is needed to prepare 200.0 ml of a 0.150 M MnCl_2 solution?

$$(0.500\text{M})(V_1) = (0.150\text{M})(200.0\text{mL})$$

$$V_1 = 60.0\text{mL}$$

- j.) 250.0ml of a 0.0100 M solution of the salt mercury(II) nitrate sits out in an open beaker. After several days the volume of solution is reduced to 200.0 ml. What is the molarity of this same solution after evaporation. If this solution is diluted to 300.0

$$\frac{(250.0 \text{ mL})(0.0100 M)}{(300.0 \text{ mL})} = M_2$$

- $$\underline{-2 - 3.0120 \text{ N}}$$

$$2.09 \text{ M} \cdot 0.4500 \text{ L} = 0.9405 \text{ mol} \cdot \frac{176 \text{ g}}{\text{mol}} = \boxed{166 \text{ g}}$$

6.00 mL of a 0.558 M ferrous gluconate solution is diluted to a new volume of 50.0 mL. What is the concentration of the diluted solution?

- $$M_2 = 0.0670 M_{\odot}$$

How many grams of potassium sulfate are needed to prepare 200.0 ml of a 0.125M solution?

- 1st

$$\text{Calculate the molar concentration of all ions in a solution that contains 82.31g of copper(II) nitrate in 912 ml of solution.}$$

$\text{Cu}(\text{NO}_3)_2 \rightarrow \text{Cu}^{+2} + 2\text{NO}_3^-$

0.481 M	$[0.481]$	$[0.963]$
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$\frac{82.31 \text{ g}}{0.9121 \text{ L}} \cdot \frac{\text{L}}{187.5 \text{ g/L}} = 0.481 \text{ M}$

- $\text{I:1} \rightarrow$ $0.912L$ $187.5g$

1:2
A student places 2.00g of salt containing chromium in some water. Later it is tested for chromium ions, however there are no chromium ions in solution. What can be concluded about the chromium salt?

THE CR SALT IS INSOLUBLE SINCE IT DIDN'T

- DISSOCIATE INTO IONS.



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